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~~solution. 8.3~~

~~Properties of Acids~~

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general properties of acids include sour taste, reactivity with metals, and ability to produce color changes in indicators.

- An acid is a compound that produces hydronium

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ions (H_3O^+) when dissolved in water. •
An indicator is any substance that changes color in the presence of an acid or base.

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STUDY. PLAY. solute.
substance that is
dissolved in a solvent
to make a solution.
solvent. substance in
which a solute is
dissolved to form a
solution. dissociation.

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The substance that is
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solute is dissolved to
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An acid produces hydronium ions in solution. Acceptable answers include hydrochloric acid, citric acid, and acetic acid. A base produces hydroxide ions in solution. Sodium

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hydroxide. 8

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Section 8.1

Formation of

Solutions (pages

228–234) This

section explains the

parts of a solution,

the processes that

occur when

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Chapter 8.4 Strength
of Acids and Bases
(pages 246–249)

This section explains
how to describe acids
and bases in terms of
both concentration
and strength.

Reading Strategy
(page 246)

Comparing and
Contrasting As you
read, complete the
diagram by

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Comparing and
contrasting acids and
bases. For more
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...

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Dissolving:
Dissociation,
Dispersion or

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Substances
can dissolve in water
by three ways: 1.
dissociation 2.
dispersion 3.
ionization

Dissociation is the
process by which an
ionic compound
separates into ions
Ex: NaCl
Dispersion is
the process by which
a compound breaks
into small pieces that

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spread throughout
the water Ex: sugar
dissolves in water
Ionization is the
process by which
neutral molecules
gain or lose electrons
(forming ions) Ex ...

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Answer Key
... solution is neutral
in the acid/base
sense. pH •To avoid
the small numbers
associated with
describing acidic and
basic solutions in
terms of mol/L, pH is
defined as

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~~Chapter 8 Acids, Bases, and Acid-Base Reactions~~

This section begins with a review of acids, followed by the following for bases: (1) it states the Arrhenius definition of base, (2) it provides you with the information necessary to identify strong and weak

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bases, and (3) it describes the changes that take place when one weak base (ammonia) dissolves in water (Figure 8.2).

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Chapter 8a Acids and
Bases July 30, 2020 1.
Definition of an

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Arrhenius (1887) acid
and base: 2.

Definition of a

Bronsted-Lowry

(1923) acid and base:

3. Write the

neutralization

reaction for: HNO_3

(aq) + NaOH (aq) 4.

Write the reaction of

vinegar (acetic acid)

and baking soda

(sodium

bicarbonate): 5. Write

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the reaction of
limestone (CaCO_3)
and HCl . 6.

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Bases.doc...~~

Chapter 8 vocabulary
(Solutions, Acids, and
Bases) 8.1- Formation
of Solutions
Ionization The
process in which
neutral molecules

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gain or lose electrons.
Solute A substance
whose particles are
dissolved in a
solution. Kool-Aid
Dispersion The
breaking into the
small pieces that

~~Chapter 8 vocabulary
(Solutions, Acids, and
Bases) by ...~~

Base: A compound
that produces

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hydroxide ions when
dissolve in water; a
proton acceptor:

Neutrization: A

chemical reaction

between and acid

and a base: Salt: An

ionic compound

formed when an acid

reacts with a base:

pH: A measure of the

hyronium ion

concentration of a

solution: Buffer: A

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and Bases chapter of
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associated with acids,
bases and solutions.
Each of these...

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times. Chemistry.
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Edit... Which

methods will distinguish between equimolar solutions of a strong base and a strong acid? I. Add magnesium to each solution and look for the formation of gas bubbles II. Add aqueous sodium hydroxide ...

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Chapter 8, Solutions,
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Particles that are

dissolved in a

solution. The solution

in which a solute will

dissolve in. The

process in which an

ionic compound

separates into ions as

it dissolves. Breaking

down into small

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Quia Chapter 8,
Solutions, Acids, and
Bases

pieces that spread
out through the
water.

Chapter 8,
Solutions, Acids, and
Bases

In Chapter 8.6, we
defined acids as
substances that
dissolve in water to
produce H^+ ions,
whereas bases were
defined as

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Substances that dissolve in water to produce OH^- ions. In fact, this is only one possible set of definitions.

~~Chapter 8.07: Acid
Base Reactions—
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Solution: The
reaction of the acid +
base gives a product
of salt + water, which~~

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Chapter 8
is considered as
neutralization
reaction. Examples:
NaOH + HCl → NaCl +
H₂O

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the year ' s entire
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